

**UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS**

**GCE Advanced Subsidiary Level and GCE Advanced Level**

**MARK SCHEME for the October/November 2011 question paper  
for the guidance of teachers**

**9701 CHEMISTRY**

**9701/23**

Paper 2 (AS Structured Questions), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- Cambridge will not enter into discussions or correspondence in connection with these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2011 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.

Page 2	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – October/November 2011	9701	23

- 1 (a) same proton number/atomic number (1)  
different mass number/nucleon number (1) [2]

(b)  $A_r = \frac{(32 \times 95.00) + (33 \times 0.77) + (34 \times 4.23)}{100}$  (1)

$$= \frac{3040 + 25.41 + 143.82}{100} = \frac{3209.23}{100}$$

which gives  $A_r = 32.09$  (1) [2]

(c)

	number of		
isotopes	protons	neutrons	electrons
$^{213}\text{Po}$	84	129	84
$^{232}\text{Th}$	90	142	90

allow **one mark** for each correct column  
if there are no 'column' marks,  
allow **maximum one mark** for a correct row

(3 × 1) [3]

- (d) (i) nucleon no. is 228 (1)  
proton no. is 88 (1)

(ii) Ra **not** radium (1) [3]

**[Total: 10]**

2 (a) (i) mass of C =  $\frac{12 \times 1.32}{44} = 0.36\text{g}$  (1)

$$n(\text{C}) = \frac{0.36}{12} = 0.03$$
 (1)

(ii) mass of H =  $\frac{2 \times 0.54}{18} = 0.06\text{ g}$  (1)

$$n(\text{H}) = \frac{0.06}{1} = 0.06$$
 (1)

- (iii) yes **because** 0.03 mol of C are combined with 0.06 mol of H **or**  
C : H ratio is 1 : 2 **or**  
empirical formula is CH<sub>2</sub> (1) [5]

Page 3	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – October/November 2011	9701	23

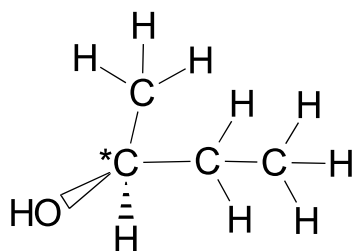
(b) (i)  $C : H : O = \frac{64.86}{12} : \frac{13.50}{1} : \frac{21.64}{16}$  (1)

= 5.41: 13.50 : 1.35

= 4 : 10 : 1

gives  $C_4H_{10}O$  (1)

(ii)

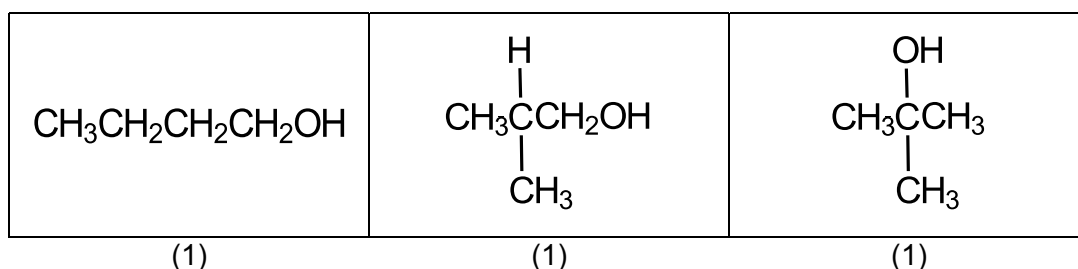


correct compound **and** correct chiral C\* (1)

correct mirror object/ mirror

image relationship in 3D (1)

(iii)



[7]

[Total: 12]

3 (a)  $C(g) \rightarrow C^+(g) + e$  (1)  
 correct equation (1)  
 correct state symbols (1) [2]

(b) (i) **Na and Mg**  
 Mg has greater nuclear charge/more protons than Na (1)

in both atoms, the 3s electrons are in the same orbital/  
 same energy level/same shell (1)

(ii) **Mg and Al**  
 in Al outermost electron is in 3p rather than 3s (1)

3p electron is at higher energy **or**  
 is further away/is more shielded from nucleus (1)

Page 4	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – October/November 2011	9701	23

- (iii) **He and Ne**  
both He and Ne have the highest nuclear charges in their Period (1)
- (iv) **He, Ne, and Ar**  
going down the group,  
valence/outer shell electrons are farther from the nucleus (1)  
there is greater shielding (1)  
attraction between valence electrons and nucleus is less **or**  
effective nuclear charge is less (1) [8]
- (c) (i) **from Na to Cl**  
increased nuclear charge/nuclear attraction (1)
- (ii) cation has fewer electrons than atom **or**  
cation has lost outer electrons **or**  
cation has fewer shells (1)  
but cation has same nuclear charge as atom **or**  
proton number is the same (1) [3]

3 (d) ignore any state symbols

MgO(s)	+	NaOH(aq)	→	NO REACTION	(1)		
MgO(s)	+	2HCl(aq)	→	MgCl <sub>2</sub> + H <sub>2</sub> O	(1)		
Al <sub>2</sub> O <sub>3</sub> (s)	+	2NaOH(aq)	+	3H <sub>2</sub> O(l)	→	2NaAl(OH) <sub>4</sub> <b>or</b>	
Al <sub>2</sub> O <sub>3</sub> (s)	+	2NaOH(aq)	+	H <sub>2</sub> O(l)	→	2NaAlO <sub>2</sub> + 2H <sub>2</sub> O <b>or</b>	(1)
Al <sub>2</sub> O <sub>3</sub> (s)	+	6NaOH(aq)	+	3H <sub>2</sub> O(l)	→	2Na <sub>3</sub> Al(OH) <sub>6</sub>	
Al <sub>2</sub> O <sub>3</sub> (s)	+	6HCl(aq)	→	2AlCl <sub>3</sub> + 3H <sub>2</sub> O <b>or</b>	(1)		
Al <sub>2</sub> O <sub>3</sub> (s)	+	6HCl(aq)	→	Al <sub>2</sub> Cl <sub>6</sub> + 3H <sub>2</sub> O			
SO <sub>2</sub> (g)	+	NaOH(aq)	→	NaHSO <sub>3</sub> <b>or</b>	(1)		
SO <sub>2</sub> (g)	+	2NaOH(aq)	→	Na <sub>2</sub> SO <sub>3</sub> + H <sub>2</sub> O			
SO <sub>2</sub> (g)	+	HCl(aq)	→	NO REACTION	(1)	[6]	

[Total: 19]

- 4 (a) (i) C<sub>2</sub>H<sub>5</sub>O (1)
- (ii)  (1) [2]

Page 5	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – October/November 2011	9701	23

- (b) (i) functional group isomerism  
or structural isomerism (1)

do **not** allow 'functional isomerism' or positional isomerism

(ii)

compound	type of isomerism
<b>P</b>	<i>cis-trans</i> or geometrical
<b>T</b>	optical

(1 + 1) [3]

- (c) (i) dehydration/elimination (1)

- (ii) conc.  $\text{H}_2\text{SO}_4$  /  $\text{P}_4\text{O}_{10}$  /  $\text{Al}_2\text{O}_3$  /  $\text{H}_3\text{PO}_4$  / pumice (1)

- (iii)  $\text{CH}_2=\text{CHCH}=\text{CH}_2$

allow  $\text{CH}_2=\text{C}=\text{CHCH}_3$  (1) [3]

- (d) (i)  $\text{CH}_3\text{CH}_2\text{CH}(\text{OH})\text{CH}_2\text{CH}_3$  (1)

- (ii) steam with  $\text{H}_3\text{PO}_4$  catalyst or  
conc.  $\text{H}_2\text{SO}_4$  then water (1 + 1)

only allow condition mark if reagent mark has been given

- (iii)  $\text{Cr}_2\text{O}_7^{2-} / \text{H}^+$  or  
 $\text{MnO}_4^- / \text{H}^+$  (1) [4]

**[Total: 12]**

- 5 (a) **V** is  $\text{HCHO}$  (1) [1]

- (b) (i) ester (1)

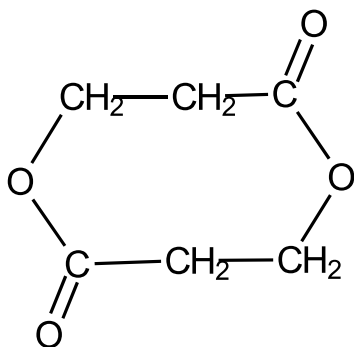
- (ii) **W** is  $\text{HCO}_2\text{CH}_3$  (1) [2]

- (c) (i) **X** is  $\text{HOCH}_2\text{CH}_2\text{CO}_2\text{H}$  (1)

- (ii) **Y** is  $\text{HO}_2\text{CCH}_2\text{CO}_2\text{H}$  (1) [2]

Page 6	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – October/November 2011	9701	23

(d) (i) Z is



(1)

(ii) esterification or  
dehydration or  
elimination or  
condensation

(1) [2]

[Total: 7]