

**UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS**

**GCE Advanced Subsidiary Level and GCE Advanced Level**

**MARK SCHEME for the October/November 2010 question paper  
for the guidance of teachers**

**9701 CHEMISTRY**

**9701/22**

Paper 2 (AS Structured Questions), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- CIE will not enter into discussions or correspondence in connection with these mark schemes.

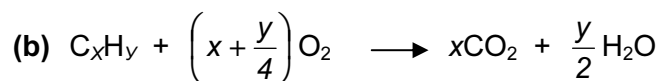
CIE is publishing the mark schemes for the October/November 2010 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.

Page 2	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE A/AS LEVEL – October/November 2010	9701	22

1 (a) the actual number of atoms of each element present (1)

in one molecule of a compound (1)

[2]



$xCO_2$  (1)

$\frac{y}{2}H_2O$  (1)

[2]

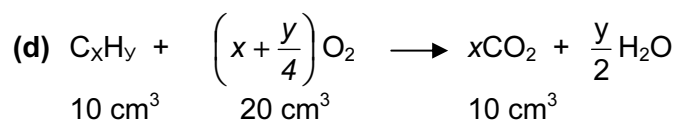
(c) (i) oxygen/ $O_2$  (1)

(ii) carbon dioxide/ $CO_2$  (1)

(iii)  $10\text{ cm}^3$  (1)

(iv)  $20\text{ cm}^3$  (1)

[4]



1 mol of  $C_xH_y$  gives 1 mol of  $CO_2$

whence  $x = 1$  (1)

1 mol of  $C_xH_y$  reacts with 2 mol of  $O_2$

whence  $\left(x + \frac{y}{4}\right) = 2$

and  $y = 4$  (1)

molecular formula is  $CH_4$  (1)

[3]

[Total: 11]

Page 3	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE A/AS LEVEL – October/November 2010	9701	22

2 (a)  $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$  (1) [1]

(b) temperature between 300 and 550°C (1)

correct explanation of effect of temperature on rate of formation of  $\text{NH}_3$  **or** on position of equilibrium (1)

catalyst of iron **or** iron oxide (1)

to speed up reaction **or** to reduce  $E_a$  (1) [4]

(c) manufacture of  $\text{HNO}_3$

**or** explosives

**or** nylon

**or** as a cleaning agent

**or** as a refrigerant (1) [1]

(d) fertiliser in rivers causes excessive growth of aquatic plants/algae (1)

when plants and algae die  $\text{O}_2$  is used up/fish or aquatic life die (1) [2]

(e) (i) CO by incomplete combustion of the hydrocarbon fuel (1)

NO by reaction between  $\text{N}_2$  and  $\text{O}_2$  in the engine (1)

(ii) CO toxic/effect on haemoglobin (1)

NO toxic/formation of acid rain (1) [4]

(f) (i) platinum/Pt – allow palladium/Pd **or** rhodium/Rh (1)

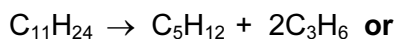
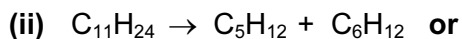
(ii)  $2\text{CO} + 2\text{NO} \rightarrow 2\text{CO}_2 + \text{N}_2$  (1) [2]

[Total: 14]

Page 4	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE A/AS LEVEL – October/November 2010	9701	22

- 3 (a) (i) a compound which contains **only** carbon and hydrogen (1)
- (ii) separation of compounds by their boiling points (1) [2]

- (b) (i) high temperature **and** high pressure (1)
- high temperature **and** catalyst (1)



- (c) (i)

$CH_3CH_2CH_2CH_2CH_3$	$CH_3CH_2CH(CH_3)CH_3$	$  \begin{array}{c}  CH_3 \\    \\  CH_3CCH_3 \\    \\  CH_3  \end{array}  $
isomer <b>B</b>	isomer <b>C</b>	isomer <b>D</b>
(1)	(1)	(1)

- (ii) the straight chain isomer (isomer **B** above) (1)

it has the greatest van der Waals' forces (1)

because unbranched molecules have greater area of contact/  
can pack more closely together (1)

[6]

- (d) enthalpy change when 1 mol of a substance (1)

is burnt in an excess of oxygen/air under standard conditions  
**or** is completely combusted under standard conditions (1)

[2]

Page 5	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE A/AS LEVEL – October/November 2010	9701	22

(e) (i) heat released =  $m c \delta T = 200 \times 4.18 \times 27.5$  (1)

$$= 22990 \text{ J} = 23.0 \text{ kJ} \text{ (1)}$$

(ii) 23.0 kJ produced from 0.47 g of **E**

$$2059 \text{ kJ produced from } \frac{0.47 \times 2059}{23.0} \text{ g of E (1)}$$

$$= 42.08 \text{ g of E (1)}$$

allow ecf in (i) or (ii) on candidate's expressions

[4]

(f)  $C_3H_6 = 42$

**E** is  $C_3H_6$

for ecf, **E** must be unsaturated and be no larger than  $C_5$  (1)

[1]

**[Total: 18]**

4 (a) reaction 1

reagent NaOH/KOH (1)

solvent  $H_2O$ /water/aqueous (1)

reaction 2

reagent  $NH_3$ /ammonia (1)

solvent ethanol/ $C_2H_5OH$ /alcohol (1)

reaction 3

reagent NaOH/KOH (1)

solvent ethanol/ $C_2H_5OH$ /alcohol (1)

[6]

(b) with  $CH_3CH_2CH_2CH_2I$  rate would be faster (1)

C-I bond is weaker than C-Br bond (1)

C-I bond energy is  $240 \text{ kJ mol}^{-1}$ , C-Br bond energy is  $280 \text{ kJ mol}^{-1}$   
data **must** be quoted for this mark (1)

[3]

(c) non-toxic

non-flammable

volatile/low bp

unreactive (any 2)

[2]

<b>Page 6</b>	<b>Mark Scheme: Teachers' version</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>GCE A/AS LEVEL – October/November 2010</b>	<b>9701</b>	<b>22</b>

(d) (i) when a covalent bond breaks the two electrons in the bond are shared between the two atoms (1)

(ii)  $CCl_2F_2 \rightarrow CClF_2 + Cl$  (as minimum)

allow  $CCl_2F + F$  (1)

[2]

(e) they are flammable (1)

[1]

**[Total: 14]**

5 (a) NaBr/sodium bromide

[1]

(b)  $Br_2$ /bromine or  $SO_2$ /sulfur dioxide

[1]

(c) concentrated sulfuric acid is an oxidising agent  
or  
phosphoric(V) acid is **not** an oxidising agent

[1]

**[Total: 3]**