



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
General Certificate of Education  
Advanced Subsidiary Level and Advanced Level

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**CHEMISTRY**

**9701/11**

Paper 1 Multiple Choice

**May/June 2013**

**1 hour**

Additional Materials: Multiple Choice Answer Sheet  
Soft clean eraser  
Soft pencil (type B or HB is recommended)  
Data Booklet



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**READ THESE INSTRUCTIONS FIRST**

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

**DO NOT WRITE IN ANY BARCODES.**

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

**Read the instructions on the Answer Sheet very carefully.**

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

Electronic calculators may be used.

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This document consists of **14** printed pages and **2** blank pages.



## Section A

For each question there are four possible answers, **A**, **B**, **C**, and **D**. Choose the **one** you consider to be correct.

- 1 Solutions containing chlorate(I) ions are used as household bleaches and disinfectants. These solutions decompose on heating as shown.



Which oxidation state is shown by chlorine in each of these three ions?

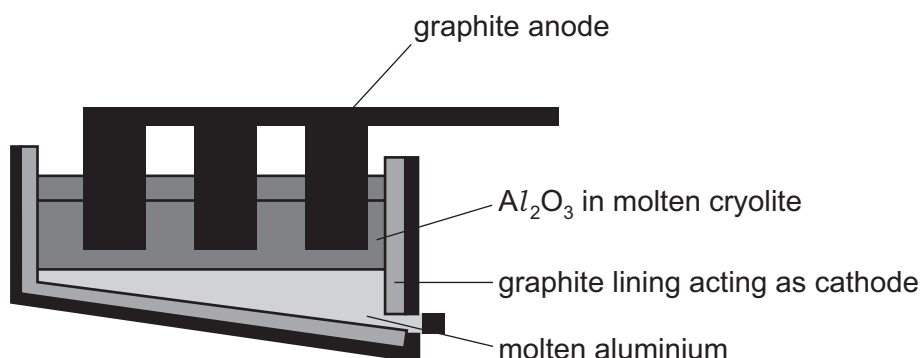
	$\text{ClO}^-$	$\text{ClO}_3^-$	$\text{Cl}^-$
<b>A</b>	+1	+3	-1
<b>B</b>	-1	+3	+1
<b>C</b>	+1	+5	-1
<b>D</b>	-1	+5	+1

- 2 A mixture of  $10\text{ cm}^3$  of methane and  $10\text{ cm}^3$  of ethane was sparked with an excess of oxygen. After cooling to room temperature, the residual gas was passed through aqueous potassium hydroxide.

All gas volumes were measured at the same temperature and pressure.

What volume of gas was absorbed by the alkali?

- A**  $15\text{ cm}^3$       **B**  $20\text{ cm}^3$       **C**  $30\text{ cm}^3$       **D**  $40\text{ cm}^3$
- 3 The diagram shows an electrolytic cell for the extraction of aluminium.



Which statement is correct?

- A** Aluminium ions are oxidised in this process.  
**B** Aluminium is liberated at the anode by the reaction  $\text{Al}^{3+} + 3\text{e}^- \rightarrow \text{Al}$ .  
**C** Cryolite is purified aluminium oxide.  
**D** The graphite anode burns away.

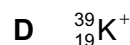
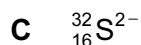
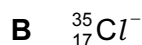
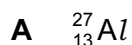
- 4 Use of the Data Booklet is relevant to this question.

The elements radon (Rn), francium (Fr) and radium (Ra) have proton numbers 86, 87 and 88 respectively.

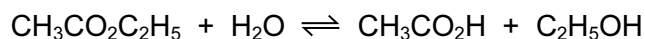
What is the order of their first ionisation energies?

	least endothermic	—————→	most endothermic
<b>A</b>	Fr	Ra	Rn
<b>B</b>	Fr	Rn	Ra
<b>C</b>	Ra	Fr	Rn
<b>D</b>	Rn	Ra	Fr

- 5 In which species are the numbers of protons, neutrons and electrons all different?



- 6 An experiment is set up to measure the rate of hydrolysis of ethyl ethanoate.

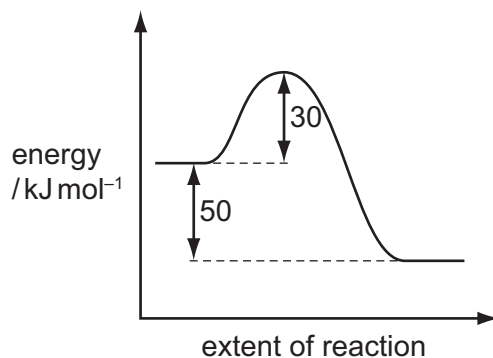


The hydrolysis is found to be slow in neutral aqueous solution but it proceeds at a measurable rate when the solution is acidified with hydrochloric acid.

What is the function of the hydrochloric acid?

- A** to dissolve the ethyl ethanoate
- B** to ensure that the reaction reaches equilibrium
- C** to increase the reaction rate by catalytic action
- D** to suppress ionisation of the ethanoic acid formed

- 7 The reaction pathway for a reversible reaction is shown below.



Which statement is correct?

- A** The activation energy of the reverse reaction is  $+80 \text{ kJ mol}^{-1}$ .
- B** The enthalpy change for the forward reaction is  $+30 \text{ kJ mol}^{-1}$ .
- C** The enthalpy change for the forward reaction is  $+50 \text{ kJ mol}^{-1}$ .
- D** The enthalpy change for the reverse reaction is  $+30 \text{ kJ mol}^{-1}$ .
- 8 Why does the rate of a gaseous reaction increase when the pressure is increased at a constant temperature?
- A** More particles have energy that exceeds the activation energy.
- B** The particles have more space in which to move.
- C** The particles move faster.
- D** There are more frequent collisions between particles.
- 9 Which would behave the **least** like an ideal gas at room temperature?
- A** carbon dioxide
- B** helium
- C** hydrogen
- D** nitrogen
- 10 The general gas equation can be used to calculate the  $M_r$  value of a gas.
- For a sample of a gas of mass  $m$  g, which expression will give the value of  $M_r$ ?
- A**  $M_r = \frac{mpV}{RT}$     **B**  $M_r = \frac{pVRT}{m}$     **C**  $M_r = \frac{mRT}{pV}$     **D**  $M_r = \frac{pV}{mRT}$

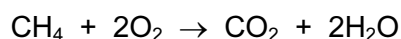
- 11 A solution of  $\text{Sn}^{2+}$  ions will reduce an acidified solution of  $\text{MnO}_4^-$  ions to  $\text{Mn}^{2+}$  ions. The  $\text{Sn}^{2+}$  ions are oxidised to  $\text{Sn}^{4+}$  ions in this reaction.

How many moles of  $\text{Mn}^{2+}$  ions are formed when a solution containing 9.5 g of  $\text{SnCl}_2$  ( $M_r$ : 190) is added to an excess of acidified  $\text{KMnO}_4$  solution?

- A 0.010                      B 0.020                      C 0.050                      D 0.125

- 12 Use of the Data Booklet is relevant to this question.

This question should be answered using bond enthalpy data. The equation for the complete combustion of methane is given below.



What is the enthalpy change of combustion of methane?

- A  $-1530 \text{ kJ mol}^{-1}$   
 B  $-1184 \text{ kJ mol}^{-1}$   
 C  $-770 \text{ kJ mol}^{-1}$   
 D  $-688 \text{ kJ mol}^{-1}$

- 13 In which row of the table are all statements comparing the compounds of magnesium and barium correct?

	solubility of hydroxides		solubility of sulfates	
	solubility of magnesium hydroxide	solubility of barium hydroxide	solubility of magnesium sulfate	solubility of barium sulfate
A	higher	lower	higher	lower
B	higher	lower	lower	higher
C	lower	higher	higher	lower
D	lower	higher	lower	higher

- 14 What happens when iodine solution is added to a solution of sodium bromide?

- A A reaction occurs without changes in oxidation state.  
 B Bromide ions are oxidised, iodine atoms are reduced.  
 C Bromide ions are reduced, iodine atoms are oxidised.  
 D No reaction occurs.



18 Sulfur trioxide is manufactured from sulfur dioxide and oxygen, using the Contact process.

Which condition affects the value of the equilibrium constant,  $K_c$ ?

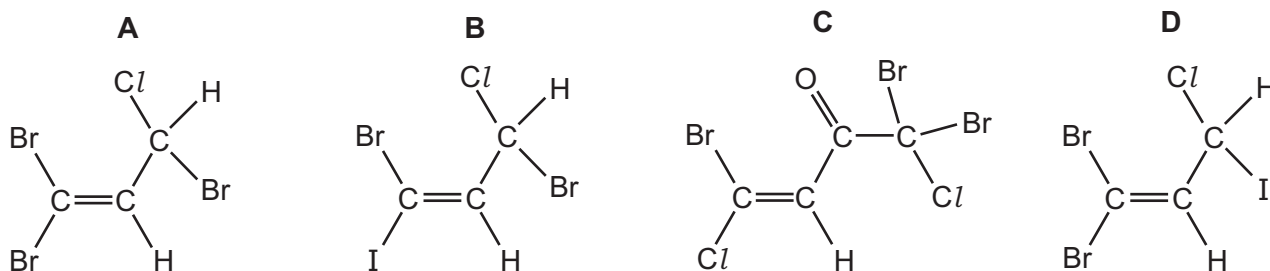
- A adjusting the temperature
- B increasing the pressure
- C removing  $\text{SO}_3$  from the equilibrium mixture
- D using a catalyst

19 Which reagent, when mixed and heated with ammonium sulfate, liberates ammonia?

- A aqueous bromine
- B dilute hydrochloric acid
- C limewater
- D potassium dichromate(VI) in acidic solution

20 The following compounds are found in the seaweed *Asparagopsis taxiformis*.

Which compound could show **both** *cis-trans* isomerism and optical isomerism?



21 Lactic acid (2-hydroxypropanoic acid),  $\text{CH}_3\text{CH}(\text{OH})\text{CO}_2\text{H}$ , is found in sour milk.

Which reaction could occur with lactic acid?

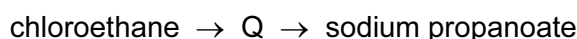
- A  $\text{CH}_3\text{CH}(\text{OH})\text{CO}_2\text{H} + \text{CH}_3\text{OH} \rightarrow \text{CH}_3\text{CH}(\text{OCH}_3)\text{CO}_2\text{H} + \text{H}_2\text{O}$
- B  $\text{CH}_3\text{CH}(\text{OH})\text{CO}_2\text{H} + \text{HCO}_2\text{H} \rightarrow \text{CH}_3\text{CH}(\text{O}_2\text{CH})\text{CO}_2\text{H} + \text{H}_2\text{O}$
- C  $\text{CH}_3\text{CH}(\text{OH})\text{CO}_2\text{H} + \text{NaHCO}_3 \rightarrow \text{CH}_3\text{CH}(\text{ONa})\text{CO}_2\text{H} + \text{H}_2\text{O} + \text{CO}_2$
- D  $\text{CH}_3\text{CH}(\text{OH})\text{CO}_2\text{H} + \text{Cl}_2 \rightarrow \text{CH}_3\text{CH}(\text{Cl})\text{CO}_2\text{H} + \text{HOCl}$

22 Bromine reacts with ethene to form 1,2-dibromoethane.

What is the correct description of the organic intermediate in this reaction?

- A It has a negative charge.
- B It is a free radical.
- C It is a nucleophile.
- D It is an electrophile.

23 Chloroethane can be used to make sodium propanoate.



The intermediate, Q, is hydrolysed with boiling aqueous sodium hydroxide, to give sodium propanoate.

Which reagent would produce the intermediate, Q, from chloroethane?

- A concentrated ammonia solution
- B dilute sulfuric acid
- C hydrogen cyanide
- D potassium cyanide

24 Aqueous sodium hydroxide reacts with 1-bromopropane to give propan-1-ol.

How should the first step in the mechanism be described?

- A by a curly arrow from a lone pair on the  $\text{OH}^-$  ion to the  $\text{C}^{\delta+}$  atom of 1-bromopropane
- B by a curly arrow from the  $\text{C}^{\delta+}$  atom of 1-bromopropane to the  $\text{OH}^-$  ion
- C by a curly arrow from the C–Br bond to the C atom
- D by the homolytic fission of the C–Br bond

25 Pentane,  $\text{C}_5\text{H}_{12}$ , is reacted with chlorine in the presence of ultraviolet light. A compound R is found in the products. R has molecular formula  $\text{C}_5\text{H}_{10}\text{Cl}_2$ . Each molecule of R contains **one** chiral carbon atom.

Which two atoms of the pentane chain could be bonded to chlorine atoms in this isomer?

- A 1 and 3
- B 1 and 5
- C 2 and 3
- D 2 and 4



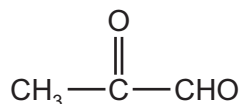
26 Use of the Data Booklet is relevant to this question.

2.30 g of ethanol were mixed with an excess of aqueous acidified potassium dichromate(VI). The reaction mixture was then boiled under reflux for one hour. The desired organic product was then collected by distillation. The yield of product was 60.0%.

What mass of product was collected?

- A 1.32 g      B 1.38 g      C 1.80 g      D 3.20 g

27 Burnt sugar has a characteristic smell caused partly by the following compound.

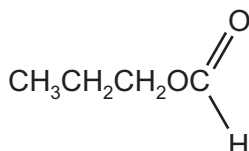


This compound contains two functional groups.

Which reagent will react with **only one** of the functional groups?

- A acidified potassium dichromate(VI)  
 B 2,4-dinitrophenylhydrazine  
 C hydrogen cyanide  
 D sodium hydroxide

28 The structural formula of a compound **X** is shown below.

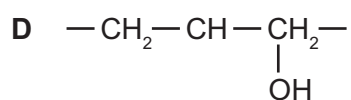
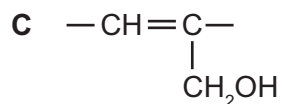
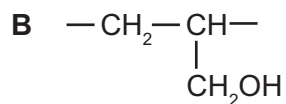
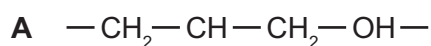


What is the name of compound **X** and how does its boiling point compare with that of butanoic acid?

	name of <b>X</b>	boiling point of <b>X</b>
<b>A</b>	methyl propanoate	higher
<b>B</b>	methyl propanoate	lower
<b>C</b>	propyl methanoate	higher
<b>D</b>	propyl methanoate	lower

- 29 Synthetic resins, plasticisers and many other chemicals can be made by polymerisation of a variety of monomers including prop-2-en-1-ol,  $\text{CH}_2=\text{CHCH}_2\text{OH}$ .

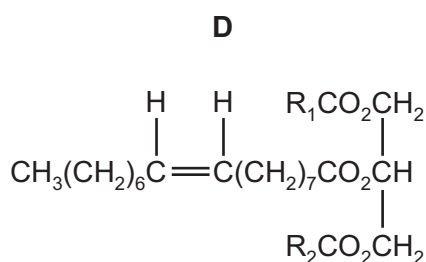
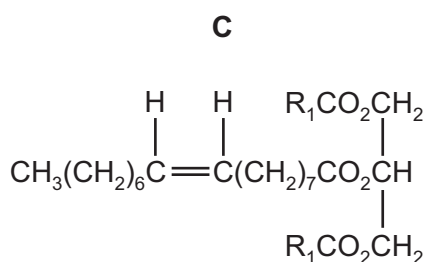
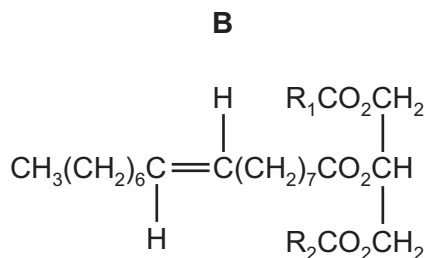
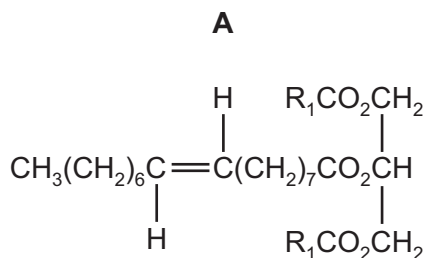
Which structure represents the repeat unit in poly(prop-2-en-1-ol)?



- 30 Some vegetable oils contain 'trans fats' that are associated with undesirable increases in the amount of cholesterol in the blood.

In the diagrams below,  $\text{R}_1$  and  $\text{R}_2$  are different hydrocarbon chains.

Which diagram correctly illustrates an optically active 'trans fat'?



## Section B

For each of the questions in this section, one or more of the three numbered statements 1 to 3 may be correct.

Decide whether each of the statements is or is not correct (you may find it helpful to put a tick against the statements that you consider to be correct).

The responses **A** to **D** should be selected on the basis of

<b>A</b>	<b>B</b>	<b>C</b>	<b>D</b>
1, 2 and 3 are correct	1 and 2 only are correct	2 and 3 only are correct	1 only is correct

No other combination of statements is used as a correct response.

**31** *Use of the Data Booklet is relevant to this question.*

Free-radicals play an important part in reactions involving the destruction of the ozone layer and the substitution of alkanes by chlorine.

Some free-radicals contain two unpaired electrons. Such species are called diradicals.

Which species are diradicals?

- 1 O
- 2 Cl
- 3 CH<sub>3</sub>

**32** The Group II metals have higher melting points than the Group I metals.

Which factors could contribute towards the higher melting points?

- 1 There are smaller interatomic distances in the metallic lattices of the Group II metals.
- 2 More electrons are available from each Group II metal atom for bonding the atom into the metallic lattice.
- 3 Group II metals have a higher first ionisation energy than the corresponding Group I metal.

**33** Valence shell electron pair repulsion theory should be used to answer this question.

Which species are trigonal planar?

- 1 BH<sub>3</sub>
- 2 CH<sub>3</sub><sup>+</sup>
- 3 PH<sub>3</sub>

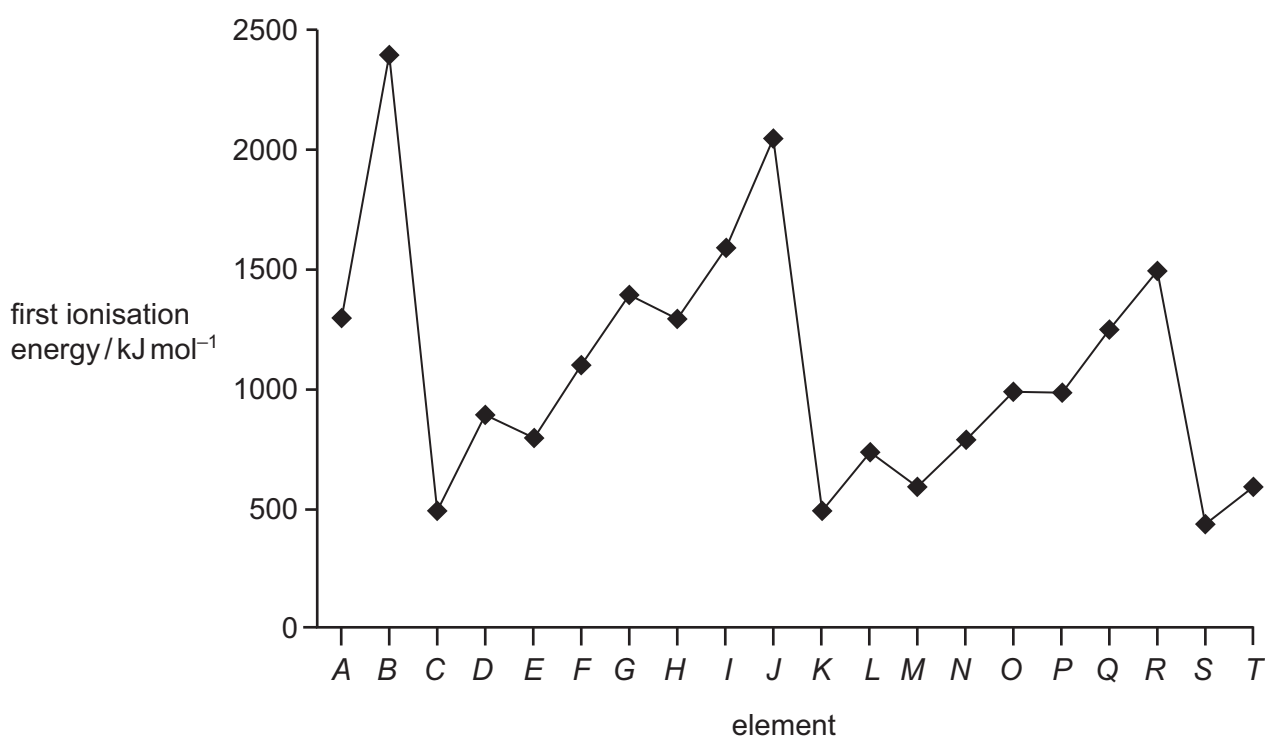
The responses **A** to **D** should be selected on the basis of

<b>A</b>	<b>B</b>	<b>C</b>	<b>D</b>
1, 2 and 3 are correct	1 and 2 only are correct	2 and 3 only are correct	1 only is correct

No other combination of statements is used as a correct response.

- 34** The first ionisation energies of twenty successive elements in the Periodic Table are represented in the graph.

The letters given are not the normal symbols for these elements.



Which statements about this graph are correct?

- 1 Elements *B*, *J* and *R* are in Group 0 of the Periodic Table.
- 2 Atoms of elements *D* and *L* contain two electrons in their outer shells.
- 3 Atoms of elements *G* and *O* contain a half-filled p subshell.

- 35 Solids **W**, **X**, **Y** and **Z** are compounds of two different Group II metals. Some of their applications are described below.

Compound **W** is used as a refractory lining material in kilns.

Compound **X** is used as a building material. It can also be heated in a kiln to form compound **Y**. When **Y** is hydrated, it forms compound **Z** which is used agriculturally to treat soils.

Which statements about these compounds are correct?

- 1 More acid is neutralised by 1 g of **W** than by 1 g of **X**.
  - 2 The metallic element in **W** reacts with water more quickly than the metallic element in **Y**.
  - 3 Adding **Z** to a soil decreases the pH of the soil.
- 36 When a red-hot platinum wire is plunged into a test tube of hydrogen iodide, the gas is decomposed into its elements. If the experiment is repeated with hydrogen chloride, no change occurs.

Which factors contribute to this behaviour?

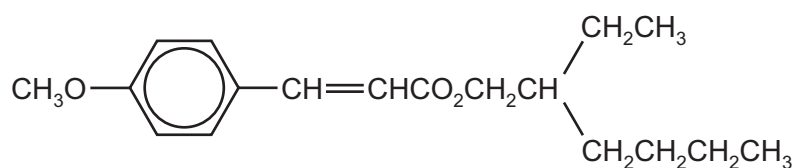
- 1 the strength of the hydrogen-halogen bond
  - 2 the size of the halogen atom
  - 3 the standard enthalpy of formation,  $\Delta H_f^\ominus$ , of each of the products of decomposition
- 37 Which molecules would be present in the mixture produced by the photochemical chlorination of methane?
- 1 hydrogen
  - 2 hydrogen chloride
  - 3 dichloromethane
- 38 In which reactions is the organic compound oxidised by the given reagent?
- 1  $\text{CH}_3\text{CH}_2\text{CHO}$  + Fehling's reagent
  - 2  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CHO}$  + Tollens' reagent
  - 3  $\text{CH}_3\text{CHO}$  + 2,4-dinitrophenylhydrazine reagent

The responses **A** to **D** should be selected on the basis of

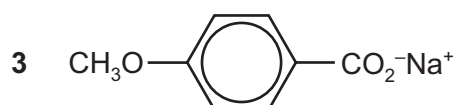
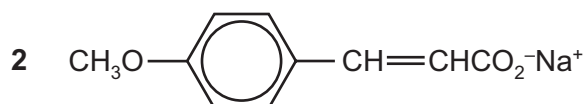
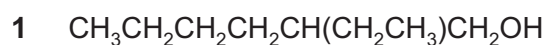
<b>A</b>	<b>B</b>	<b>C</b>	<b>D</b>
1, 2 and 3 are correct	1 and 2 only are correct	2 and 3 only are correct	1 only is correct

No other combination of statements is used as a correct response.

39 A sun protection cream contains the following ester as its active ingredient.



Which substances are present in the products of its hydrolysis by aqueous sodium hydroxide?



40 Which reagents, when used in an excess, can be used to make sodium lactate,  $\text{CH}_3\text{CH}(\text{OH})\text{CO}_2\text{Na}$ , from lactic acid,  $\text{CH}_3\text{CH}(\text{OH})\text{CO}_2\text{H}$ ?

1 Na

2  $\text{NaHCO}_3$

3 NaOH



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